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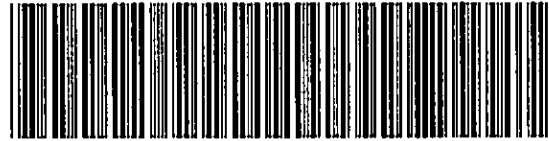
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2021 DEC -3 AM 10:19  
SEC. OF STATE

**ARTICLES OF AMENDMENT  
TO  
ARTICLES OF ORGANIZATION  
OF**

ESTUDIO MARSHALL & PARTNERS LLC

(Name of the Limited Liability Company as it now appears on our records.)  
(A Florida Limited Liability Company)

The Articles of Organization for this Limited Liability Company were filed on 03/14/2016 and assigned Florida document number L16000052233.

This amendment is submitted to amend the following:

**A. If amending name, enter the new name of the limited liability company here:**

The new name must be distinguishable and contain the words "Limited Liability Company," the designation "LLC" or the abbreviation "L.L.C."

Enter new principal offices address, if applicable:

(Principal office address MUST BE A STREET ADDRESS)

Enter new mailing address, if applicable:

(Mailing address MAY BE A POST OFFICE BOX)

**B. If amending the registered agent and/or registered office address on our records, enter the name of the new registered agent and/or the new registered office address here:**

Name of New Registered Agent:

New Registered Office Address:

Enter Florida street address

Florida

City

2021 DEC 19

Zip Code

**New Registered Agent's Signature, if changing Registered Agent:**

Whereby accept the appointment as registered agent and agree to act in this capacity. I further agree to comply with the provisions of all statutes relative to the proper and complete performance of my duties, and I am familiar with and accept the obligations of my position as registered agent as provided for in Chapter 605, F.S. Or, if this document is being filed to merely reflect a change in the registered office address, I hereby confirm that the limited liability company has been notified in writing of this change.

If Changing Registered Agent, Signature of New Registered Agent

If amending Authorized Person(s) authorized to manage, enter the title, name, and address of each person being added or removed from our records:

MGR = Manager  
AMBR = Authorized Member

<u>Title</u>	<u>Name</u>	<u>Address</u>	<u>Type of Action</u>
AMBR	LUCAS MARSHALL	11281 SW 9TH MANOR, DAVIE , FL 33325	<input checked="" type="checkbox"/> Add
			<input type="checkbox"/> Remove
			<input type="checkbox"/> Change
AMBR	CAMILA MARSHALL	11281 SW 9TH MANOR, DAVIE, FL 33325	<input checked="" type="checkbox"/> Add
			<input type="checkbox"/> Remove
			<input type="checkbox"/> Change
			<input type="checkbox"/> Add
			<input type="checkbox"/> Remove
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			<input type="checkbox"/> Remove
			<input type="checkbox"/> Change

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Q. No.	Q. Text	Ans.
Q. 11	Write the chemical equation for the reaction of sodium metal with water.	$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$
Q. 12	Write the chemical equation for the reaction of calcium metal with water.	$\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
Q. 13	Write the chemical equation for the reaction of magnesium metal with water.	$\text{Mg} + 2\text{H}_2\text{O} \rightarrow \text{Mg(OH)}_2 + \text{H}_2$
Q. 14	Write the chemical equation for the reaction of zinc metal with water.	$\text{Zn} + 2\text{H}_2\text{O} \rightarrow \text{Zn(OH)}_2 + \text{H}_2$
Q. 15	Write the chemical equation for the reaction of iron metal with water.	$\text{Fe} + 2\text{H}_2\text{O} \rightarrow \text{Fe(OH)}_2 + \text{H}_2$
Q. 16	Write the chemical equation for the reaction of aluminium metal with water.	$2\text{Al} + 6\text{H}_2\text{O} \rightarrow 2\text{Al(OH)}_3 + 3\text{H}_2$
Q. 17	Write the chemical equation for the reaction of sodium metal with oxygen.	$4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$
Q. 18	Write the chemical equation for the reaction of calcium metal with oxygen.	$2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$
Q. 19	Write the chemical equation for the reaction of magnesium metal with oxygen.	$2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
Q. 20	Write the chemical equation for the reaction of zinc metal with oxygen.	$2\text{Zn} + \text{O}_2 \rightarrow 2\text{ZnO}$
Q. 21	Write the chemical equation for the reaction of iron metal with oxygen.	$4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
Q. 22	Write the chemical equation for the reaction of aluminium metal with oxygen.	$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
Q. 23	Write the chemical equation for the reaction of sodium metal with chlorine.	$2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
Q. 24	Write the chemical equation for the reaction of calcium metal with chlorine.	$\text{Ca} + \text{Cl}_2 \rightarrow \text{CaCl}_2$
Q. 25	Write the chemical equation for the reaction of magnesium metal with chlorine.	$\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
Q. 26	Write the chemical equation for the reaction of zinc metal with chlorine.	$\text{Zn} + \text{Cl}_2 \rightarrow \text{ZnCl}_2$
Q. 27	Write the chemical equation for the reaction of iron metal with chlorine.	$\text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_2$
Q. 28	Write the chemical equation for the reaction of aluminium metal with chlorine.	$2\text{Al} + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3$
Q. 29	Write the chemical equation for the reaction of sodium metal with sulphur.	$2\text{Na} + \text{S} \rightarrow \text{Na}_2\text{S}$
Q. 30	Write the chemical equation for the reaction of calcium metal with sulphur.	$\text{Ca} + \text{S} \rightarrow \text{CaS}$
Q. 31	Write the chemical equation for the reaction of magnesium metal with sulphur.	$\text{Mg} + \text{S} \rightarrow \text{MgS}$
Q. 32	Write the chemical equation for the reaction of zinc metal with sulphur.	$\text{Zn} + \text{S} \rightarrow \text{ZnS}$
Q. 33	Write the chemical equation for the reaction of iron metal with sulphur.	$\text{Fe} + \text{S} \rightarrow \text{FeS}$
Q. 34	Write the chemical equation for the reaction of aluminium metal with sulphur.	$2\text{Al} + 3\text{S} \rightarrow \text{Al}_2\text{S}_3$
Q. 35	Write the chemical equation for the reaction of sodium metal with phosphorus.	$6\text{Na} + \text{P}_4 \rightarrow 2\text{Na}_3\text{P}$
Q. 36	Write the chemical equation for the reaction of calcium metal with phosphorus.	$3\text{Ca} + \text{P}_4 \rightarrow \text{Ca}_3\text{P}_2$
Q. 37	Write the chemical equation for the reaction of magnesium metal with phosphorus.	$3\text{Mg} + \text{P}_4 \rightarrow \text{Mg}_3\text{P}_2$
Q. 38	Write the chemical equation for the reaction of zinc metal with phosphorus.	$3\text{Zn} + \text{P}_4 \rightarrow \text{Zn}_3\text{P}_2$
Q. 39	Write the chemical equation for the reaction of iron metal with phosphorus.	$3\text{Fe} + \text{P}_4 \rightarrow \text{Fe}_3\text{P}_2$
Q. 40	Write the chemical equation for the reaction of aluminium metal with phosphorus.	$4\text{Al} + 3\text{P}_4 \rightarrow \text{Al}_4\text{P}_3$
Q. 41	Write the chemical equation for the reaction of sodium metal with nitrogen.	$6\text{Na} + \text{N}_2 \rightarrow 2\text{Na}_3\text{N}$
Q. 42	Write the chemical equation for the reaction of calcium metal with nitrogen.	$3\text{Ca} + \text{N}_2 \rightarrow \text{Ca}_3\text{N}_2$
Q. 43	Write the chemical equation for the reaction of magnesium metal with nitrogen.	$3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$
Q. 44	Write the chemical equation for the reaction of zinc metal with nitrogen.	$3\text{Zn} + \text{N}_2 \rightarrow \text{Zn}_3\text{N}_2$
Q. 45	Write the chemical equation for the reaction of iron metal with nitrogen.	$3\text{Fe} + \text{N}_2 \rightarrow \text{Fe}_3\text{N}_2$
Q. 46	Write the chemical equation for the reaction of aluminium metal with nitrogen.	$4\text{Al} + 3\text{N}_2 \rightarrow \text{Al}_4\text{N}_3$
Q. 47	Write the chemical equation for the reaction of sodium metal with carbon.	$2\text{Na} + \text{C} \rightarrow \text{Na}_2\text{C}$
Q. 48	Write the chemical equation for the reaction of calcium metal with carbon.	$\text{Ca} + \text{C} \rightarrow \text{CaC}$
Q. 49	Write the chemical equation for the reaction of magnesium metal with carbon.	$\text{Mg} + \text{C} \rightarrow \text{MgC}$
Q. 50	Write the chemical equation for the reaction of zinc metal with carbon.	$\text{Zn} + \text{C} \rightarrow \text{ZnC}$
Q. 51	Write the chemical equation for the reaction of iron metal with carbon.	$\text{Fe} + \text{C} \rightarrow \text{FeC}$
Q. 52	Write the chemical equation for the reaction of aluminium metal with carbon.	$2\text{Al} + 3\text{C} \rightarrow \text{Al}_2\text{C}_3$
Q. 53	Write the chemical equation for the reaction of sodium metal with hydrogen.	$2\text{Na} + \text{H}_2 \rightarrow 2\text{NaH}$
Q. 54	Write the chemical equation for the reaction of calcium metal with hydrogen.	$\text{Ca} + \text{H}_2 \rightarrow \text{CaH}_2$
Q. 55	Write the chemical equation for the reaction of magnesium metal with hydrogen.	$\text{Mg} + \text{H}_2 \rightarrow \text{MgH}_2$
Q. 56	Write the chemical equation for the reaction of zinc metal with hydrogen.	$\text{Zn} + \text{H}_2 \rightarrow \text{ZnH}_2$
Q. 57	Write the chemical equation for the reaction of iron metal with hydrogen.	$\text{Fe} + \text{H}_2 \rightarrow \text{FeH}_2$
Q. 58	Write the chemical equation for the reaction of aluminium metal with hydrogen.	$2\text{Al} + 3\text{H}_2 \rightarrow 2\text{AlH}_3$
Q. 59	Write the chemical equation for the reaction of sodium metal with fluorine.	$2\text{Na} + \text{F}_2 \rightarrow 2\text{NaF}$
Q. 60	Write the chemical equation for the reaction of calcium metal with fluorine.	$\text{Ca} + \text{F}_2 \rightarrow \text{CaF}_2$
Q. 61	Write the chemical equation for the reaction of magnesium metal with fluorine.	$\text{Mg} + \text{F}_2 \rightarrow \text{MgF}_2$
Q. 62	Write the chemical equation for the reaction of zinc metal with fluorine.	$\text{Zn} + \text{F}_2 \rightarrow \text{ZnF}_2$
Q. 63	Write the chemical equation for the reaction of iron metal with fluorine.	$\text{Fe} + \text{F}_2 \rightarrow \text{FeF}_2$
Q. 64	Write the chemical equation for the reaction of aluminium metal with fluorine.	$2\text{Al} + 3\text{F}_2 \rightarrow 2\text{AlF}_3$
Q. 65	Write the chemical equation for the reaction of sodium metal with bromine.	$2\text{Na} + \text{Br}_2 \rightarrow 2\text{NaBr}$
Q. 66	Write the chemical equation for the reaction of calcium metal with bromine.	$\text{Ca} + \text{Br}_2 \rightarrow \text{CaBr}_2$
Q. 67	Write the chemical equation for the reaction of magnesium metal with bromine.	$\text{Mg} + \text{Br}_2 \rightarrow \text{MgBr}_2$
Q. 68	Write the chemical equation for the reaction of zinc metal with bromine.	$\text{Zn} + \text{Br}_2 \rightarrow \text{ZnBr}_2$
Q. 69	Write the chemical equation for the reaction of iron metal with bromine.	$\text{Fe} + \text{Br}_2 \rightarrow \text{FeBr}_2$

(If an effective date is listed, the date must be specific and cannot be prior to date of filing or more than 90 days after filing.) Pursuant to 605.0207 (3)(b)

If the record specifies a delayed effective date, but not an effective time, at 12:01 a.m. on the earlier of: (b) The 90th day after the record is filed.

Signature of a member or authorized representative

Typed or printed name of signee